

UNIVERSITY OF TORONTO
STUDIES

PAPERS FROM THE CHEMICAL
LABORATORIES

No. 74: SOME ESTERS OF ARSENIOS
J. F. MACKEY AND R. A. GORTNER

(REPRINTED FROM THE JOURNAL OF THE CHEMICAL SOCIETY)

THE UNIVERSITY LIBRARY: PUBLISHED BY
THE LIBRARIAN, 1908

NTO

EMICAL

ENIOUS ACID, BY W. R. LANG,

EMICAL SOCIETY, VOL. 93)

PUBLISHED BY

Unive
COMM

Chairman: ROBERT A

PROFESSOR W

PROFESSOR W

PROFESSOR A.

PROFESSOR J.

PROFESSOR R.

PROFESSOR G

General Editor:

University of Toronto Studies

COMMITTEE OF MANAGEMENT

ROBERT ALEXANDER FALCONER, M.A., Litt.D., LL.D., D.D.
President of the University

PROFESSOR W. J. ALEXANDER, Ph.D.

PROFESSOR W. H. ELLIS, M.A., M.B.

PROFESSOR A. KIRSCHMANN, Ph.D.

PROFESSOR J. J. MACKENZIE, B.A.

PROFESSOR R. RAMSAY WRIGHT, M.A., B.Sc., LL.D.

PROFESSOR GEORGE M. WRONG, M.A.

Librarian: H. H. LANGTON, M.A.

Librarian of the University

CXXX.—*Some Esters of Arsenious*

By WILLIAM ROBERT LANG, JOHN FRANCIS
ROSS AITKEN GORTNER.

J. M. CRAFTS (*Bull. Soc. chim.*, 1870, [ii], 14, 9) has pointed out the existence of compounds of arsenic with alcohol, and has shown that no esters of arsenious or of arsenic acid had been known up to the year 1870, and described a method by which ethyl, methyl, and amyl arsenates; this consisted in heating the corresponding iodides in sealed tubes at 100° with arsenate. The esters were purified by washing with water and distilling under diminished pressure. They are liquids.

Arsenious Acid.

FRANCIS MACKEY, and
TNER.

ii], 14, 99), in referring to
with alcohol radicles, pointed
onic acid had been prepared
ethod by which he obtained
is consisted in heating the
at 100° with normal silver
washing with ether and dis-
are liquids of high boiling

point, above 200° , and suffer partial decomposition if distilled under atmospheric pressure; in the case of the amyl arsenate so much so that he found it impossible to obtain it pure even when distilled in a vacuum. The addition of water to them causes immediate and complete decomposition into arsenic acid and the alcohol. Crafts determined their composition by weighing the precipitated arsenic acid as the magnesium salt and the carbon and hydrogen by combustion, and gave them the formula R_3AsO_4 . For the arsenites of the alkyl radicles he used three methods, a sealed tube being employed: (1) the interaction of arsenious oxide and ethyl silicate at 200° , (2) the interaction of ethyl iodide and silver arsenite at 150° ; and (3) the interaction of arsenic bromide and sodium ethoxide, which last method he considered the best, although in it a secondary reaction between the ester formed and the sodium ethoxide causes a condition of equilibrium to be set up, and no more ester is formed.

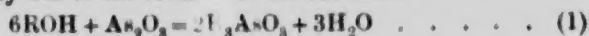
It is to be noted that even if disodium hydrogen arsenite is employed the resulting arsenite is always the trialkyl salt. Crafts states that arsenious oxide and alcohol do not react when heated in a sealed tube, nor does arsenious oxide either with ether alone or with the addition of ethyl acetate, even when in contact for twenty hours.

These esters prepared by Crafts are specified by Auger (*Compt. rend.*, 1902, **134**, 238) as the only ones known, and he refers in a later paper (*Compt. rend.*, 1906, **43**, 907) to Crafts' inability to obtain any ester by heating the alcohols and oxides of arsenic in a sealed tube. Auger bases his conviction that an ester is produced on the fact that arsenious oxide is volatile in alcohol vapour, although not in water vapour, and concludes that this cannot be accounted for unless there is an ester formed.* Also he succeeded in preparing an ester from glycerol and arsenic oxide, and from ethyl, methyl, isopropyl, isobutyl, and isoamyl alcohols and arsenious oxide. He heated the alcohols with crystalline arsenious oxide for some hours, evidently in a sealed tube, as the temperatures quoted are far above the boiling points of the alcohols, and gives 6.5 per cent. as the extreme limit of esterification in the case of methyl alcohol, and 1.2 per cent. in that of ethyl alcohol. Calculating from his figures, the yield of ester for the higher alcohols appears to be 2.62 per cent., 0.25 per cent., 1.00 per cent., and 0.63 per cent. for propyl, isopropyl, isobutyl, and isoamyl alcohols respectively.

So far as can be gathered from the paper, Auger takes no account of the extreme solubility, as we have found, of arsenious oxide in the ester. These small yields and the difficulty found by Crafts in

* We have passed the vapour of methyl alcohol over arsenious oxide heated in a long tube, and the ether obtained differs entirely from methyl arsenite.

obtaining any results from the action of the acid on the alcohol are evidently due to the reversal of the reaction:



Anger states that he removed the water as it was formed by fractional distillation (in the case of *isobutyl* and *isoamyl* esters), or by passing the mixture of alcohol and water vapour over calcium carbide placed in an adapter, thus removing the water and allowing the alcohol to drop back into the flask in which the reaction takes place. An excellent yield, how much is not stated, of propyl, *n*-butyl and *isobutyl* arsenites was thus obtained. Phenol, in the same way, yielded an ester hitherto only prepared by the action of sodium phenoxide on arsenic trichloride.

The nature of the reaction expressed by equation (1) and the properties of the ester clearly show that only by removing the water as it is formed can one expect to get a large yield. Our work has been done in two ways, namely, by heating the mixture in a vessel to which is attached a modified Soxhlet tube and condenser, the Soxhlet containing anhydrous copper sulphate, and by adding this dehydrating agent directly to the mixture. A comparison of the yields obtained is given in the sequel. By these two methods yields of from 14.8 per cent. (cold) to 58.6 per cent. (hot) have been obtained with aliphatic alcohols, and with the phenols as high as 90 per cent. Experiments have also been made, with very encouraging results, on benzyl alcohol and on esters of hydroxy-acids. We hope to show that the use of such a dehydrating agent as copper sulphate in a Soxhlet tube will allow of the preparation of a series of esters from the oxides and possibly sulphides of arsenic, antimony, tin, and perhaps bismuth with compounds containing hydroxyl, where the boiling points of such esters are higher than that of water, as, indeed, is the case with all of these; also with anhydrous copper sulphate in contact with the substances themselves. In many preparations, too, of esters generally and allied products, the formation of water in the reaction produces an equilibrium resulting in a very low yield. It is expected that by removing this water, as is done in the instances described in the present paper, the yields will be greatly increased in many commercial processes.

EXPERIMENTAL.

Heating with Inverted Condenser only.—Weighed quantities of propyl, *isobutyl*, and *isoamyl* alcohols were mixed with excess of arsenious oxide and heated with a direct flame for forty-five minutes in a flask to which a reflux condenser was attached. The clear liquid obtained was poured off from the excess of arsenious oxide and fractionated under diminished pressure. The esters so purified were

analysed and the yield compared with that calculated from equation (1).

Our method of ascertaining the composition of the esters differs from that used by Auger.* The ester was decomposed with water, forming arsenious oxide and liberating the alcohol. The arsenious oxide was dissolved in sodium carbonate and titrated with *N*/10 iodine (a preliminary test having shown that the alcohol did not react with iodine). By this means the quantity of arsenic in the ester was determined.

isoAmyl Arsenite.—1.4 Grams were decomposed with 10 c.c. of water, sufficient sodium carbonate added to dissolve the arsenious oxide, and the whole diluted to 100 c.c.; 25 c.c. of this required 17 c.c. of standard iodine for oxidation, corresponding with 22.4 per cent. of arsenic. In 50 grams of the ester, the alcohol was determined and found to weigh 39 grams, or 78 per cent. *isoButyl* and *propyl* arsenites were also analysed in a similar way:

Ester.	Per cent. As found.	Per cent. As calculated as (RO) ₃ As.
Methyl.....	44.5	44.6
Ethyl	—	—
Propyl	29.6	29.7
<i>isoButyl</i>	25.7	25.5
Amyl	22.2	22.3
<i>isoAmyl</i>	22.4	22.3

Expressing these as salts of arsenious acid, the formulæ become $(C_5H_{11})_3AsO_3$, $(C_4H_9)_3AsO_3$, and $(C_3H_7)_3AsO_3$.

Propyl arsenite is a yellow, mobile liquid boiling at 216° and decomposing very readily on addition of water.

isoButyl arsenite is a deep yellow, mobile liquid of specific gravity 1.069; it decomposes rapidly in presence of water into *isobutyl* alcohol and arsenious oxide; under 760 mm. pressure it decomposes at 242°, and boils at 157° under 30 mm. pressure.

isoAmyl arsenite is a yellow liquid of specific gravity 1.050; it boils at 185° under 30 mm. pressure, and under atmospheric pressure it decomposes at 284°; in the presence of water, it decomposes into *isoamyl* alcohol and arsenious oxide.

Heating with Soxhlet Attachment and Anhydrous Copper Sulphate.—To eliminate the water as it is formed and so increase the yield of ester, the conditions of the experiments were modified. One hundred and sixty grams of *isoamyl* alcohol were added to 70 grams of

* Auger determined the amount of arsenic in the cooled liquid obtained (a mixture of the ester and alcohol) by means of iodine. The difference between this and the amount of arsenious oxide that would dissolve in the same quantity of alcohol in the cold was taken as representing the amount of ester formed (*Compt. rend.*, 1906, 143, 908). No account was taken of the solubility of arsenious oxide in the esters themselves.

arsenious oxide in a 300 c.c. flask. An ordinary Soxhlet tube was connected with the flask, and a condenser, fitted with a calcium chloride tube, was attached to the Soxhlet, which contained a large filter paper filled with anhydrous copper sulphate. The liquid in the flask was heated to boiling, and the water formed was absorbed by the anhydrous copper sulphate, which turned blue as soon as the first drops of condensed liquid fell on it. Heating was stopped after forty-five minutes, when small rings of a compound of high boiling point began to form on the sides of the flask; after cooling, the clear product was filtered and fractionated under diminished pressure. 110.25 grams of ester were obtained, being a yield of 58.6 per cent. as compared with 13.22 per cent. obtained without the use of a dehydrating agent.

In similar experiments with *isobutyl* and *propyl* alcohols, yields of 56.25 per cent. and 56.75 per cent. respectively were obtained, as compared with 11.09 and 8.79 per cent. when no dehydrating agent was employed.

Use of Dehydrating Agent in the Cold.—The esterification of these three alcohols by arsenious oxide can be carried out even in the cold if a dehydrating agent is used; 120 grams of the alcohol were shaken in a stoppered bottle at room temperature with 60 grams of arsenious oxide and 70 grams of anhydrous copper sulphate for three days. The following yields were obtained: *isoamyl* arsenite, 17.2 per cent.; *isobutyl* arsenite, 15.8 per cent.; *propyl* arsenite, 14.8 per cent.

The following table indicates the comparative yields (per cent.) by different methods:

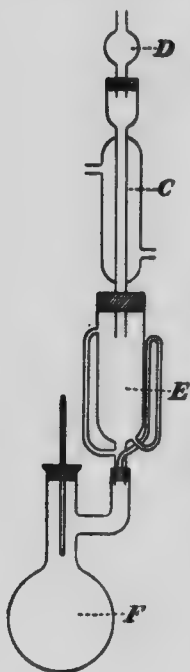
Ester.	Auger (in sealed tube).	Reflux condenser.	With anhydrous copper sulphate:	
			in Soxhlet.	in the cold.
Methyl	8.5	—	33.8*	—
Ethyl	1.2	—	4.5*	—
Propyl	2.62	8.79	56.75	14.8
<i>isoButyl</i>	0.25	11.09	56.25	15.8
Trimethylcarbinol	—	—	54.27	—
Amyl	1.00	—	74.00	—
<i>isoAmyl</i> ..	0.63	13.22	58.62	17.2

* The Soxhlet was not used, but the anhydrous copper sulphate was placed with the alcohol and arsenious oxide in a flask fitted with a reflux condenser.

A probable explanation of the low yield in the case of the ethyl arsenite lies in the fact that ethyl alcohol itself has a strong affinity for water which, for small quantities of water present, cannot be overcome by the anhydrous copper sulphate. This is readily shown by adding a drop of water to 5 c.c. of ethyl alcohol in which a little anhydrous copper sulphate has been placed. The copper sulphate is apparently not affected, whilst in the case of the higher alcohols a blue colour appears almost instantly, showing the hydration of the

copper sulphate. Methyl alcohol does not respond to this test as readily as do the higher alcohols, but it responds more readily than does ethyl alcohol. The yield in each case, using the copper sulphate, probably bears a direct relation to the respective affinities of the alcohol and of the copper sulphate for the water of the reaction. If a substance could be used which had a greater dehydrating power and at the same time was insoluble in alcohol, a higher yield would doubtless be obtained. Calcium carbide (Auger, *Compt. rend.*, 1906, 143, 908) is not such a material: we have compared the yields

obtained by replacing the anhydrous copper sulphate in the Soxhlet with fresh calcium carbide, using various alcohols, and the highest yield was in the case of *isoamyl* alcohol where 40.2 per cent. of ester was obtained after two hours' heating.



Esters obtained with Phenol and its Homologues.

For the esterification of these with arsenious oxide a side-necked flask was used the neck of which was bent upwards and attached to the Soxhlet, leaving the mouth of the flask free to receive a thermometer (see figure).

Phenyl Arsenite.—Weighed quantities of phenol (140 grams) and arsenious oxide (80 gram-) were heated together, and in all cases it was observed that the reaction began at 100°, the mixture boiling violently at that temperature. The thermometer gradually rose to 133°, when a thick cloud formed in the flask. The temperature remained constant at that point for a few minutes and then gradually rose to a maximum, where it was kept for about five minutes: the heating was then stopped and the contents of the flask allowed to cool. The mixture of phenol, ester, and arsenious

oxide, the latter of which is very soluble in the ester, was shaken with benzene, causing the precipitation of the arsenious oxide dissolved in the ester. The mixture was then filtered and the benzene solution of the ester fractionated under diminished pressure. The benzene distilled at 20°, the phenol at 69°, and the ester at 305°. The ester, purified by redistillation, weighed 105 grams, representing a yield of 60 per cent. *Phenyl arsenite* is a deep yellow, viscous liquid with a specific gravity of 1.59; it freezes at 31°, and boils at 305° under a pressure of 30 mm. It dissolves readily in methyl alcohol, benzene, ethyl acetate, or chloroform, and decomposes on addition of water, but not

so readily as the fatty arsenites. It was found on analysis to correspond with the formula $(C_6H_5O)_3As$, or $(C_6H_5)_3AsO_3$.

o-, *m*-, and *p*-Tolyl Arsenites.—One hundred grams of each of the cresols were heated with arsenious oxide for thirty minutes in the apparatus previously described, the arsenious oxide being separated by benzene. The liquid was then fractionated, and the esters, after having been purified and analysed as in the case of the phenyl arsenite, were found to correspond with the formulæ $(C_6H_4Me)_3As$, or $(C_6H_4Me)_3AsO_3$.

Naphthyl Arsenite.—One hundred grams of naphthol were heated with 30 grams of arsenious oxide for thirty minutes and the arsenious oxide in the ester separated by means of benzene. As yet, however, the pure ester has not been isolated.

Benzyl Arsenite.—One hundred grams of benzyl alcohol were heated with 35 grams of arsenious oxide for thirty minutes. The maximum temperature was 240° . The clear liquid was decanted and 75 c.c. of benzene added, the arsenious oxide collected, and the filtrate fractionated under 30 mm. pressure. The *benzyl arsenite* distilled at 285° , suffering partial decomposition. Although in the case of the fatty alcohols a drying agent in the Soxhlet was essential to absorb the water formed and to allow only the alcohols themselves to drop into the mixture, no such device was necessary with the phenols, as practically no phenol or ester ever found its way further than the side tube of the Soxhlet. The water formed remained in this apparatus, there being never sufficient volume of liquid produced to cause the syphon attachment to come into play.

To ascertain if it was possible to obtain these esters without the removal of the water formed during the reaction, the same quantities of phenol, benzyl alcohol, *o*-, *m*-, and *p*-cresol respectively, as were used in the previous experiments, were heated with arsenious oxide in a flask fitted with a reflux condenser only, but in no case was any arsenite formed in quantities sufficient to enable it to be isolated.

Method of Analysis for Esters of Phenol and its Homologues.—To determine the composition of the esters, it was found necessary to distil them several times in order to remove all traces of arsenious oxide which is readily soluble in them, in some cases to the extent of 30 per cent. About 3 grams of the ester were decomposed by 5 c.c. of water and 10 c.c. of potassium hydroxide (containing 700 grams of potassium hydroxide per litre), and the whole diluted to 500 c.c. Two samples of 10 c.c. each were taken for analysis; to one sample standard iodine was added in excess, shown by the formation of a precipitate of tri-iodophenol and the appearance of a clear yellow solution, the temperature being kept at 65° . The mixture was then cooled, acidified with sulphuric acid, and diluted to 500 c.c. with

water. The excess of iodine in 100 c.c. of this was titrated against standard thiosulphate, using starch as indicator. The quantity of thiosulphate necessary, multiplied by five, represents the quantity of iodine used. This quantity, subtracted from the original amount of iodine added, gave the amount necessary to change both the phenol to tri-iodophenol and the arsenious oxide to arsenic oxide. To determine the quantity of arsenious oxide present in the ester, an excess of standard potassium dichromate was added to the second sample and this excess was determined with standard ferrous sulphate. The ratio between the dichromate and iodine being known, the number of c.c. of iodine equivalent to the amount of dichromate used was found and subtracted from the total iodine obtained in the previous determination, the difference being the amount combined with the phenol. From these data the relative quantities of arsenious oxide and phenol, formed by decomposing the ester, were obtained, and the composition of the ester thus determined.

To test the accuracy of this method, estimations were made with weighed quantities of (a) phenol, (b) arsenious oxide, (c) a mixture of these.

(a) To 0.8826 gram of phenol, 5 c.c. of potassium hydroxide (700 grams per litre) were added and the whole diluted to 250 c.c. . . (1)

Ten c.c. of this solution were found to be equivalent to 25.75 c.c. of standard iodine.

(b) To 0.5 gram of arsenious oxide, 5 c.c. of the same potassium hydroxide were added and diluted to 250 c.c. (2)

Ten c.c. of this solution were found to be equivalent to 6.5 c.c. of standard iodine.

(c) To 10 c.c. of solution (1), 10 c.c. of solution (2) were added and found to require 32.30 c.c. of standard iodine.

Thus 10 c.c. of solution (1) required 25.75 c.c. of iodine, and 10 c.c. of solution (2) required 6.50 c.c. of iodine; in all 32.25 c.c. as compared with 32.30 c.c. when mixed, a difference which is well within the limits of experimental error.

In order to see if arsenious oxide can be determined accurately in the presence of phenol by means of dichromate, experiments similar to those made with iodine were carried out with it. Ten c.c. of solution (2) required 6.07 c.c. of dichromate, a mixture of 10 c.c. of (2) with 10 c.c. of (1) required 6.09 c.c. of dichromate, or an error of less than 0.3 per cent.

Properties of the Arsenites prepared from Phenol and its Homologues.

	Phenyl arsenite.	Benzyl arsenite.	<i>o</i> -Tolyl arsenite.	<i>m</i> -Tolyl arsenite.	<i>p</i> -Tolyl arsenite.
Yield (per cent.)					
(a) with Soxhlet	60	100	96	94	95
(b) without Soxhlet	nil	nil	nil	nil	nil
Specific gravity	1.59	1.43	—	1.45	1.46
Refractive index *	—	1.572	—	—	—
Boiling point † under 30 mm..	305°	255°	—	346°	—
Freezing point	-31°	-36°	—	—	—
Colour	yellow	blue	dark brown	dark brown	brown

* The blanks indicate that the refractive index is greater than 1.62098, the limit of the prism used.

† Where the boiling point is not given it is above 360°.

All these esters are soluble in methyl and ethyl alcohols, ether, benzene, ethyl acetate, or chloroform, and are decomposed at once by water.

This method of esterification is being carried out with arsenious oxide and the dihydric and trihydric phenols, but the quantitative results are not yet ready. A successful attempt has also been made to form similar esters with hydroxy-acids, methyl salicylate being heated with arsenious oxide. The products of the reaction, namely, water and an oil, were driven up into the Soxhlet where the latter decomposed, liberating arsenious oxide. The oil boils at about the same temperature as the methyl salicylate and has an almost unbearable odour. When the dehydrating agent is used in the Soxhlet it is expected the new ester will readily be separated.

Experiments have also been tried with arsenious sulphide in place of the oxide and a small yield of an ester obtained, presumably of the composition R_2AsS_2 , but the upper portions of the flask and the condenser became coated with the orange-coloured arsenious sulphide, showing that decomposition had occurred. The work is being continued.

CHEMICAL LABORATORY,
UNIVERSITY OF TORONTO.



